Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Pd: \_\_\_ Ast: \_\_\_\_\_

**Ionic Bonds Practice**

*For help, refer to the periodic table and your handout with the table of common ions.*

**Write the correct charge on the line beside each ion:**

1. F \_\_\_\_\_ 2. Be \_\_\_\_\_ 3. Na \_\_\_\_\_ 4. Ca \_\_\_\_\_ 5. O \_\_\_\_\_

6. Cl \_\_\_\_\_ 7. S \_\_\_\_\_ 8. B \_\_\_\_\_ 9. K \_\_\_\_\_ 10. I \_\_\_\_\_

**Write the correct charge on the line beside each polyatomic ion:**

11. NH4 \_\_\_\_\_ 12. CO3 \_\_\_\_\_ 13. HCO3 \_\_\_\_\_ 14. PO4 \_\_\_\_\_ 15. SO4 \_\_\_\_\_

**Write the correct formula for each ionic compound:**

16. Li+ and F- \_\_\_\_\_\_\_\_\_\_ 17. Mg2+ and F- \_\_\_\_\_\_\_\_\_\_

18. Al3+ and S2- \_\_\_\_\_\_\_\_\_\_ 19. K+ and Cl- \_\_\_\_\_\_\_\_\_\_

20. Ca2+ and O2- \_\_\_\_\_\_\_\_\_\_ 21. Li+ and P3- \_\_\_\_\_\_\_\_\_\_

22. K+ and Br- \_\_\_\_\_\_\_\_\_\_ 23. Na+ and PO43- \_\_\_\_\_\_\_\_\_\_

24. NH4+ and CO32- \_\_\_\_\_\_\_\_\_\_ 25. Mg2+ and SO42- \_\_\_\_\_\_\_\_\_\_

**Write the correct name next to the formula for each ionic compound:**

26. Li2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

27. CaI2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

28. K2CO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

29. (NH4)2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

30. Al(HCO3)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Write the correct formula next to the name for each ionic compound:**

31. Potassium Sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

32. Sodium Fluoride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

33. Magnesium Sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

34. Aluminum Phosphate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

35. Ammonium Bicarbonate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_