Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Pd: \_\_\_ Ast: \_\_\_\_\_

**Ionic Bonds Practice**

*For help, refer to the periodic table and your handout with the table of common ions.*

**Write the correct charge on the line beside each ion:**

1. F \_\_\_\_\_ 2. Be \_\_\_\_\_ 3. Na \_\_\_\_\_ 4. Ca \_\_\_\_\_ 5. O \_\_\_\_\_

6. Cl \_\_\_\_\_ 7. S \_\_\_\_\_ 8. B \_\_\_\_\_ 9. K \_\_\_\_\_ 10. I \_\_\_\_\_

**Write the correct charge on the line beside each polyatomic ion:**

 11. NH4 \_\_\_\_\_ 12. CO3 \_\_\_\_\_ 13. HCO3 \_\_\_\_\_ 14. PO4 \_\_\_\_\_ 15. SO4 \_\_\_\_\_

**Write the correct formula for each ionic compound:**

 16. Li+ and F- \_\_\_\_\_\_\_\_\_\_ 17. Mg2+ and F- \_\_\_\_\_\_\_\_\_\_

 18. Al3+ and S2- \_\_\_\_\_\_\_\_\_\_ 19. K+ and Cl- \_\_\_\_\_\_\_\_\_\_

 20. Ca2+ and O2- \_\_\_\_\_\_\_\_\_\_ 21. Li+ and P3- \_\_\_\_\_\_\_\_\_\_

 22. K+ and Br- \_\_\_\_\_\_\_\_\_\_ 23. Na+ and PO43- \_\_\_\_\_\_\_\_\_\_

 24. NH4+ and CO32- \_\_\_\_\_\_\_\_\_\_ 25. Mg2+ and SO42- \_\_\_\_\_\_\_\_\_\_

**Write the correct name next to the formula for each ionic compound:**

 26. Li2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 27. CaI2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 28. K2CO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 29. (NH4)2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 30. Al(HCO3)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 **Write the correct formula next to the name for each ionic compound:**

 31. Potassium Sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 32. Sodium Fluoride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 33. Magnesium Sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 34. Aluminum Phosphate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 35. Ammonium Bicarbonate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_