Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Pd: \_\_\_\_\_ Ast: \_\_\_\_\_

**Acids & Bases**

1. **THE pH SCALE**
   1. What does the pH scale measure?
   2. How is the dissociation of water related to pH?
   3. Fill in the table below to relate the concentration of ions to each type of solution:

|  |  |
| --- | --- |
| **Type of Solution** | **Relative Concentration of Ions** |
| Neutral Solutions |  |
| Acidic Solutions |  |
| Basic Solutions |  |

1. **ACIDS**
   1. Stronger acids produce more \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ( ) ions in solution.
   2. Acids are represented by pH levels of \_\_\_\_\_ to \_\_\_\_\_.
   3. List some common acids:
   4. List three common properties of acids:
2. **BASES**
   1. Stronger bases produce more \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ( ) ions in solution.
   2. Bases are represented by pH levels of \_\_\_\_\_ to \_\_\_\_\_.
   3. List some common bases:
   4. List three common properties of bases:
3. **SALTS**
   1. How are salts related to acids and bases?
   2. What are often the products of a neutralization reaction?
   3. Explain why neutralization reactions don’t always result in a pH of 7.